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Key indicators

Single-crystal X-ray study T = 100 KMean σ (C–C) = 0.001 Å Disorder in main residue R factor = 0.043 wR factor = 0.117 Data-to-parameter ratio = 34.9

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

A cocrystal of 1-(4-chlorophenyl)-3-(3,4,5trimethoxyphenyl)prop-2-en-1-one and 3-(3-chloro-4,5-dimethoxyphenyl)-1-(4chlorophenyl)prop-2-en-1-one (0.95:0.05)

In the title cocrystal, 1-(4-chlorophenyl)-3-(3,4,5-trimethoxyphenyl)prop-2-en-1-one–3-(3-chloro-4,5-dimethoxyphenyl)-1-(4-chlorophenyl)prop-2-en-1-one (19/1), $0.95C_{18}H_{17}ClO_{4}$ ·- $0.05C_{17}H_{14}Cl_2O_3$, the crystal packing is stabilized by weak intermolecular C—H···O and π - π interactions. Received 24 May 2006 Accepted 31 May 2006

Comment

Chalcone derivatives are receiving increasing interest from several disciplines because of their photochemical properties (Matsushima & Kageyama, 1985; Toda *et al.*, 1998), their structural features, and their second- and third-order nonlinear optical properties (Fichou *et al.*, 1988; Kitaoka *et al.*, 1990; Uchida *et al.*, 1998; Goto *et al.*, 1991; Patil *et al.*, 2006*a,b*; Zhang *et al.*, 1990; Zhao *et al.*, 2000). In order to obtain detailed information on the molecular conformation of the trimethoxy-substituted chalcone (I) in the solid state, its X-ray crystal structure was determined. Refinement revealed a cocrystal with (I) as the major component (95%) but with a minor component in which the methoxy substituent at C12 has been replaced by a Cl atom (II) (Fig. 1). The compounds crystallize in a centrosymmetric space group and therefore have no second-order nonlinear optical properties.



Bond lengths and angles show normal values (Allen *et al.*, 1987) and are comparable to those reported in related structures (Ng *et al.*, 2006*a*,*b*; Patil *et al.*, 2006*a*,*b*). The enone group and the two benzene rings of the chalcone are each planar, with a maximum deviation of 0.047 (1) Å from O1/C7–C9,

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Figure 1

The structure of the title compound, showing 50% probability displacement ellipsoids and the atomic numbering. The dashed lines indicate hydrogen bonds. The hollow bond is to the Cl atom of the minor cocrystal component.



Figure 2

The crystal packing of (I), viewed down the a axis. Hydrogen bonds are shown as dashed lines.

0.013 (1) Å from C1–C6, and 0.017 (1) Å from C10–C15 for atoms C8, C6 and C14, respectively.

The short H1A····H8A (2.21 Å) contact causes the C1– C6–C7 [123.28 (7)°] and C6–C7–C8 [118.37 (7)°] bond angles to deviate significantly from 120°. Similarly, the short H8A···H15A (2.17 Å) contact produces a slight widening of the C9–C10–C15 angle to [121.83 (7)°]. These result in the molecules being twisted about the C6–C7 bond for both (I) and (II). The least-squares plane through the enone group makes dihedral angles of 16.94 (3) and 13.61 (3)° with the C1– C6 and C10–C15 benzene rings, respectively. The dihedral angles between the two benzene rings is 30.28 (3)°.

The methoxy groups at C12 and C14 are almost coplanar with the attached rings, with torsion angles $C16-O2-C12-C11 = -4.1 (1)^{\circ}$ and $C18-O4-C14-C15 = 3.0 (1)^{\circ}$, while the methoxy group at C13 is twisted away from the attached ring, with a torsion angle $C17-O3-C13-C14 = 56.4 (1)^{\circ}$.

In the crystal structure, atom O1 is involved in both intraand intermolecular hydrogen bonds (Table 1). The intramolecular C17-H17C···O4 and C9-H9A···O1 interactions generate S(5) and S(6) ring motifs, respectively (Bernstein *et al.*,1995). The crystal packing is further stabilized by weak intermolecular π - π interactions, involving the C1-C6 ring at (x, y, z) and the C10–C15 ring at (1 - x, -y, 1 - z); the centroid–centroid distance is 3.8771 (5) Å.

Experimental

Compound (I) was synthesized by the Claisen-Schmidt condensation of 3,4,5-trimethoxybenzaldehyde (0.01 mol) with 4-chloroacetophenone (0.01 mol) in ethanol (60 ml) in the presence of NaOH (2 ml, 30%). After stirring for 2 h, the contents of the flask were poured into ice-cold water (250 ml) and left to stand for 24 h. The resulting crude solid was collected by filtration, dried and purified by repeated recrystallization from acetone. Crystals suitable for singlecrystal X-ray diffraction were grown by slow evaporation of an acetone solution at room temperature over a period of 7 d. Our intention was to synthesize (I), but the presence of chlorine in 3,4,5trimethoxybenzaldehyde (one of the starting materials) in the form of 3-chloro-4,5-dimethoxybenzaldehyde led to the formation of (II) as well. The presence of a Cl atom in the minor cocrystal component was consistent with the chemical analysis. Analysis found: C 64.81, H 04.98, O 19.12%; calculated for $C_{17.95}H_{16.85}Cl_{1.05}O_{3.95}$: C 64.74, H 05.10, O 18.98%.

Z = 4

 $D_x = 1.403 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation

 $\mu = 0.27 \text{ mm}^-$

 $R_{\rm int} = 0.048$

 $\theta_{\rm max} = 37.5^\circ$

T = 100.0 (1) K Block, yellow

 $0.48 \times 0.29 \times 0.20 \text{ mm}$

54905 measured reflections

8205 independent reflections

6659 reflections with $I > 2\sigma(I)$

Crystal data

0.95C₁₈H₁₇ClO₄·0.05C₁₇H₁₄Cl₂O₃ $M_r = 332.99$ Monoclinic, $P2_1/c$ a = 7.5517 (1) Å b = 15.9894 (2) Å c = 13.6914 (2) Å $\beta = 107.491$ (1)° V = 1576.76 (4) Å³

Data collection

Bruker SMART APEX2 CCD areadetector diffractometer

Absorption correction: multi-scan (SADABS; Bruker, 2005) $T_{min} = 0.829, T_{max} = 0.949$

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_o^2) + (0.0593P)^2]$
$R[F^2 > 2\sigma(F^2)] = 0.043$	+ 0.3778P]
$wR(F^2) = 0.118$	where $P = (F_0^2 + 2F_c^2)/3$
S = 1.03	$(\Delta/\sigma)_{\rm max} = 0.001$
8205 reflections	$\Delta \rho_{\rm max} = 0.60 \ {\rm e} \ {\rm \AA}^{-3}$
235 parameters	$\Delta \rho_{\rm min} = -0.29 \text{ e } \text{\AA}^{-3}$
Only H-atom coordinates refined	

Table T			
Hydrogen-bond	geometry ((Å,	°).

$D - H \cdots A$	$D-\mathrm{H}$	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdots A$
C9−H9A···O1	0.93	2.51	2.829 (1)	100
C17−H17C···O4	0.96	2.34	2.866 (1)	114
$C9-H9A\cdotsO1^{i}$	0.93	2.51	3.414 (1)	166
$C18-H18C\cdots O1^{ii}$	0.96	2.59	3.526 (1)	164

Symmetry codes: (i) -x, -y, -z + 1; (ii) -x + 1, -y, -z + 1.

H atoms were placed in calculated positions, with C—H distances in the range 0.93–0.96 Å, and refined as riding. The $U_{\rm iso}$ values were freely refined. The ratio of (I) and (II) in the cocrystal was obtained by refinement as 0.954 (3):0.046 (3). Data collection: *APEX2* (Bruker, 2005); cell refinement: *APEX2*; data reduction: *SAINT* (Bruker, 2005); program(s) used to solve structure: *SHELXTL* (Sheldrick, 1998); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*, *PARST* (Nardelli, 1995) and *PLATON* (Spek, 2003).

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References

- Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). J. Chem. Soc. Perkin Trans. 2, pp. S1–19.
- Bernstein, J., Davis, R. E., Shimoni, L. & Chamg, N.-L. (1995). Angew. Chem. Int. Ed. Engl. 34, 1555–1573.
- Bruker (2005). *APEX2* (Version 1.27), *SAINT* (Version 7.12A) and *SADABS* (Version 2004/1). Bruker AXS Inc., Madison, Wisconsin, USA.

- Fichou, D., Watanabe, T., Takeda, T., Miyata, S., Goto, Y. & Nakayama, M. (1988). Jpn. J. Appl. Phys. 27, L429–L430.
- Goto, Y., Hayashi, A., Kimura, Y. & Nakayama, M. (1991). J. Cryst. Growth, 108, 688–698.
- Kitaoka, Y., Sasaki, T., Nakai, S., Yokotani, A., Goto, Y. & Nakayama, M. (1990). Appl. Phys. Lett. 56, 2074–2076.
- Matsushima, R. & Kageyama, H. (1985). J. Chem. Soc. Perkin Trans. 2, pp. 743–748.
- Nardelli, M. (1995). J. Appl. Cryst. 28, 659.
- Ng, S. L., Shettigar, V., Razak, I. A., Fun, H. K., Patil, P. S. & Dharmaprakash, S. M. (2006a). Acta Cryst. E62, 01421–01423.
- Ng, S. L., Shettigar, V., Razak, I. A., Fun, H. K., Patil, P. S. & Dharmaprakash, S. M. (2006b). Acta Cryst. E62, o1570–o1572.
- Patil, P. S., Teh, J. B. J., Fun, H. K., Razak, I. A. & Dharmaprakash, S. M. (2006a). Acta Cryst. E62, 0896–0898.
- Patil, P. S., Teh, J. B. J., Fun, H. K., Razak, I. A. & Dharmaprakash, S. M. (2006b). Acta Cryst. E62, 01710–01712.
- Sheldrick, G. M. (1998). SHELXTL. Version 5.1. Bruker AXS Inc., Madison, Wisconsin, USA.
- Spek, A. L. (2003). J. Appl. Cryst. 36, 7-13.
- Toda, F., Tanaka, K. & Kato, M. (1998). J. Chem. Soc. Perkin Trans. 1, pp. 1315–1318.
- Uchida, T., Kozawa, K., Sakai, T., Aoki, M., Yoguchi, H., Abdureyim, A. & Watanabe, Y. (1998). Mol. Cryst. Liq. Cryst. 314, 135–140.
- Zhang, G., Kinoshita, T., Sasaki, K., Goto, Y. & Nakayam, M. (1990). J. Cryst. Growth, 100, 411–416.
- Zhao, B., Lu, W.-Q., Zhou, Z.-H. & Wu, Y. (2000). J. Mater. Chem. 10, 1513– 1517.